

1. Introduction
 The purpose of this study is to investigate the effect of the concentration of the solution on the rate of the reaction between hydrogen peroxide and potassium iodide.

2. Aim
 To determine the effect of the concentration of the solution on the rate of the reaction.

3. Variables
 Independent variable: Concentration of the solution.
 Dependent variable: Rate of the reaction.

4. Materials
 Hydrogen peroxide, Potassium iodide, Water, Conical flask, Measuring cylinder, Stopwatch.

5. Method
 1. Prepare a series of solutions with different concentrations of hydrogen peroxide.
 2. Add a fixed volume of potassium iodide solution to each of the prepared solutions.
 3. Measure the time taken for the reaction to occur (indicated by the appearance of a blue color).
 4. Repeat the experiment for each concentration.

6. Results
 The results show that as the concentration of the hydrogen peroxide solution increases, the rate of the reaction also increases.

7. Conclusion
 The concentration of the solution has a direct effect on the rate of the reaction.

8. References
 None.